Name:_____

8.4 Polar Bonds and Molecules Notetaker

Nonpolar Covalent Bonds:

Polar Covalent Bond:

Explain how electronegativity affects the charge on a molecule:

How would you symbolize that the chlorine in HCl attracts more electrons?

Which type of bond (nonpolar covalent, moderately polar covalent, very polar covalent, or ionic) will form between each of the following pairs of atoms?

a. N and H
b. F and F
c. Ca and Cl
d. Al and Cl

How does a polar bond usually affect a molecule's polarity?

Name:_____

Dipole:

Intermolecular Forces: Van der Waals Forces: dipole interactions and Dispersion Forces, and Hydrogen Bonds

Dipole Interactions:

Dispersion Forces:

Hydrogen Bonds:

Explain why the properties of covalent compounds have such a wide variety of properties.