

Covalent Bonding Notetaker

1. Ionic Bond: _____

2. The compound potassium fluoride is made up of _____ and

_____.

3. Covalent Bonds are similar to ionic bonds because _____

4. Covalent Bonds are different from ionic bonds because _____

5. The shared electron pair is called a _____

6. Chlorine forms a bond with itself. Write the formula for that molecule below.

7. Draw a model of what happens when two chlorine atoms bond together.

8. Take a minute to explain in your own words how an octet is achieved or formed when two chlorine atoms bond together.

9. Circle the Bonding Pair. It is called a _____



10. Draw a model below with the covalent bond as a dash that shows that the two chlorines have bonded together.

11. Draw a model below of what happens when two oxygens bond together.

12. In your own words explain why oxygen forms two covalent bonds(which we will call a double bond).

Naming Covalent Bonds Notetaker

13. Compound: _____

14. Molecule: _____

15. Properties of Covalent Compounds:

16. Types of Bonds:

Single Bond:

Double Bond:

Tripple Bond:

17. Diatomic Elements:

18. Use the name of the element to name the following diatomic molecules.

H₂ hydrogen

N₂ nitrogen

Cl₂ _____

O₂ _____

I₂ _____

19. Naming Covalent Compounds:

- Two nonmetals

- Name each element
- End the last element in -ide
- Add prefixes to show more than 1 atom
- Prefixes

mon	1	hexa	6
di	2	hepta	7
tri	3	octa	8
tetra	4	nona	9
penta	5	deca	10

20. Fill in the blanks to complete the following names of covalent compounds.

CO	carbon _____oxide
CO ₂	carbon _____
PCl ₃	phosphorus _____chloride
CCl ₄	carbon _____chloride
N ₂ O	_____nitrogen _____oxide

21. Learning Check!

- A. P₂O₅
- 1) phosphorus oxide
 - 2) phosphorus pentoxide
 - 3) diphosphorus pentoxide
- B. Cl₂O₇
- 1) dichlorine heptoxide
 - 2) dichlorine oxide
 - 3) chlorine heptoxide
- C. Cl₂
- 1) chlorine
 - 2) dichlorine
 - 3) dichloride